

Comparative study of soil test methods for determination of plant available potassium in Bulgarian arable soils

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Submitted 23 March 2021, received in revised form 31 July 2021

This comparative study was aimed at estimating analytical behavior of methods for determination of plant available potassium applied to Bulgarian arable soils and to reveal the relationship between the amount of extractable K. Twenty-four samples from two traditional agricultural regions in Bulgaria were studied. Soil potassium was extracted by NH₄OAc/HOAc pH 4.5 (**AA**), diluted double acid (**Mehlich 1**), CaCl₂, BaCl₂ and a modified acetate/lactate method (**ALM**) and determined by Flame AES. The factors influencing the methods accuracy were identified and uncertainty was estimated. The expanded uncertainty was (in mg K₂O (100 g dry soil)⁻¹): 0.10 (ALM), 0.64 (Mehlich 1), 0.17 (CaCl₂) and 1.1 (AA). The study revealed that the factor which mainly influence the uncertainty of the applied analytical methods for plant available potassium in soil was the calibration of Flame AES determination. The obtained results showed that extractable potassium lowered in the following order $K_{ALM} \geq K_{AA} > K_{Mehlich1} > K_{BaCl_2} > K_{CaCl_2}$. Soil potassium extracted by ALM procedure correlated with AA, BaCl₂-K, CaCl₂-K and Mehlich 1 - K at 0.05 level of significance. ALM extracted between 1.2 to 5.8 times more soil K than other methods did. The obtained results provided a base for further study on correlation between extractable K and soil fertility indices for particular soil types and climatic regions in Bulgaria.

Keywords: available potassium, extraction methods, Flame AES, uncertainty, arable soils.

INTRODUCTION

Soil diagnostic is usually based on concentration of potassium, nitrogen and phosphorous, together with pH and organic carbon determination. Some meso and microelements could also be included for more precise fertilization. The soil nutrients occur in several forms, which put the challenge to choose appropriate extractants and detection finishing of the analysis. Moreover, a multielement analysis is highly demanded to reduce significantly labor, cost and time of analysis.

Potassium exists in soils as water-soluble (soils solution), exchangeable/available, fixed/non-exchangeable and mineral potassium (incorporated in the crystal lattice of mineral and inaccessible for plants). The extracting solution contains a mixture of various reagents that reacts with soil and releases some K into the solution. Many different extracting solutions have been used worldwide due to variable nature of soil. The Good agricultural practices in Bulgaria implied for fertilization programs based on soil diagnostic by soil analysis [1]. As it has been well recognized no universal soil testing method existed. A given fertilization program should be well reasoned based on field studies in which crop yield response to K fertilizer addition was validated against the amount K extracted by a specific reagent mixture. Additionally, the instrumental technique for determination

of K in soil extract also influenced the data [2]. Therefore, fertilizer recommendations based on a soil test should be developed for a particular extractant and detection technique [3, 4].

Difficulties in comparing the results from various methods for soil analysis for K determination arose from: (1) different forms of K in soils; (2) different extraction efficiency of the reagent mixture; (3) variations in the extraction procedure within the same extractants; (4) great diversity of soil nature and thus variation of matrix composition; (5) different detection techniques used to measure K concentration in soil extract. Nowadays, research on improving the performance of soil testing laboratories and providing sustainable fertilization programs have been constantly intensified. A line of comparative studies on K extraction methods combined with modern statistical and numerical modelling tools proposed conversation between soil test results without loss of validation data [5-10]. Multielement extracting solutions aimed at decreasing analysis costs and labor were proposed [2, 8, 11-13]. The main obstacles come from the different nature of soil and specific response of crops, which implies thorough evaluation of soil testing method for specific region and crop [2, 5, 7, 9].

To correspond to the large variety of arable soil types in Bulgaria, a modified acetate-lactate containing

extractant, based on Egner-Riehm method was proposed [14]. The method showed increased buffer capacity towards high content of free carbonates commonly encountered in Bulgarian region. The extractant was proved efficient for extraction of plant available potassium and validated for a line of crops [1, 14]. A recent study showed the possibility of replacement of equilibrium potassium in soil solution with exchangeable potassium by the modified method in estimation of potassium supply and fertilization strategy [15]. The comparison of the data about the fertility status even in a given small region has been hindered as different methods for determination of plant available K were used [1, 16-19]. Moreover, data about method performance applied to Bulgarian soils could be scarcely found.

The aims of this study were: (1) to reveal the factors which influenced the accuracy of the chosen analytical methods for potassium determination and to estimate uncertainty of measurement; (2) to reveal the relationship between the amount of plant available potassium, extracted by $\text{NH}_4\text{OAc}/\text{HOAc}$, NaHCO_3 (**Olsen**), diluted double acid (Mehlich 1), CaCl_2 , BaCl_2 and modified acetate-lactate extractant; (3) to demonstrate the relationship between the obtained amount of extractable potassium for Bulgarian arable soils. To our knowledge, the analytical behavior of the modified acetate-lactate based method [14] was presented for the first time.

MATERIALS AND METHODS

Soil sampling

The soil samples used in this study were collected from two well-developed agricultural regions in Bulgaria: North-East Bulgaria ($43^\circ 16' 01.8'' \text{N } 27^\circ 00' 55.9'' \text{E}$) and South East Bulgaria Thracian valley ($42^\circ 15' 43.0'' \text{N } 26^\circ 16' 17.3'' \text{E}$). Twenty-four arable soil samples were studied, 12 from each region. All samples were collected from a depth of 0-30 cm by automatic soil sampler following the sampling plan recommended by Nikolova *et al.* [1]. In average one sample for each 6 ha, each sample was formed from 15-20 individual samples. The samples were stored in plastic bags and transported to the laboratory, where were air dried, ground and sieved through 2 mm sieve.

Determination of soil characteristics

Clay content was determined gravimetrically as a fraction below $63 \mu\text{m}$. *pH* of the soil samples was determined in water at 1 : 2 (soil-to-liquid ratio) after 1 h mixing. The samples were left to settle and the pH was measured using WTW pH-meter equipped with a combined glass electrode.

Electrical conductivity was measured according to the procedure, described in Soil standard testing method [20]: 5 g of each soil sample were extracted with 25 mL d. H_2O for 1 h. The suspension was left to settle for 10 min and the electrical conductivity of the supernatant was measured by WTW Multi 3410

electrical conductivity meter, equipped with WTW Tetra Con 925 electrode.

Cation exchange capacity (CEC) was determined according to Hendershot and Duquette [21]. Two g of accurately weighted soil samples were treated with 20.00 mL of 0.1 M BaCl_2 solution. The suspension was stirred for 2 h and filtered. The filtrate was made up to 100.0 mL with d. H_2O and subjected to ICP-OES determination of Al, Ca, K, Mg and Na. A set of calibration standards at the concentration ranges 5 - 100 $\mu\text{g L}^{-1}$ or 1 - 10 mg L^{-1} , according to sample concentrations, were prepared by appropriate dilution of a multi element standard solution ("Ultra scientific", Lot: P00332) containing 24 elements in 5% HNO_3 at concentration $100 \pm 5 \text{ mg L}^{-1}$ of each element. Each standard was scanned at least three times and a mean analytical signal for each component was calculated.

Soil organic carbon was determined after oxidation with $\text{K}_2\text{Cr}_2\text{O}_7$ by back titration with standard solution of Mohr salt (Turin's method). Accurately weighted 0.12 g – 0.15 g of soil samples were oxidized by 10.00 mL 0.0667 M $\text{K}_2\text{Cr}_2\text{O}_7$ (in 1 : 2 H_2SO_4) solution with AgSO_4 as a catalyst. The samples were homogenized and heated for 20 min in 160°C . After cooling, the samples were diluted with distilled water and titrated by 0.2 M standard solution of Mohr salt and N-phenylanthranilic acid solution as an indicator. Two blank samples (pumice stone) were prepared following the same procedure. The $\text{Fe}(\text{NH}_4)_2(\text{SO}_4)_2$ solution was standardized against $\text{K}_2\text{Cr}_2\text{O}_7$.

Mineral nitrogen (NO_3^- -N and NH_3^- -N) was determined after diluted acid extraction at 1 : 5 soil-to-liquid ratio for 5 min followed by spectrophotometric determination of nitrates by cadmium reduction method (Method 8039, Hach Lange) and ammonium by Nessler reaction (Method 8038, Hach Lange).

Available phosphorous was determined spectrophotometrically by the molybdenum blue method after extraction with modified acetate-lactate reagent. Color reagent with ascorbic acid as a reductant was prepared daily by mixing 1.056 g ascorbic acid and 200 mL of Murphy-Reilly solution, 12 g ammonium molybdate, 0.2928 g potassium antimonyl tartrate in 2 L 1.25 M H_2SO_4 . For P_2O_5 determination, 2 mL of soil extract were mixed with the color reagent and volume was made up to 25 mL in a measuring flask. After 1 h, the absorbance was measured at 880 nm against a blank sample by Hach Lange DR 3900 Spectrophotometer. Calibration curve was daily prepared.

Determination of plant available potassium in soil

Soil potassium was extracted by $\text{NH}_4\text{OAc}/\text{HOAc}$ pH 4.5 (**AA**), diluted double acid (**Mehlich 1**), CaCl_2 , BaCl_2 and a modified acetate/lactate method (**ALM**).

K extraction by the modified acetate-lactate method (ALM). A modified acetate-lactate extractant [14] contained 0.1 M calcium lactate, 0.2 M $\text{CH}_3\text{COONH}_4$ and 0.1 M HCl, pH 4.2. All reagents were of p.a. grade,

produced by Sigma-Aldrich. The pH of the solution was controlled potentiometrically and if necessary pH was corrected with aqueous ammonia or acetic acid. Soil K extraction was carried out for 1 hour at soil-to-liquid ratio 1 : 25 (2 g accurately weighted soil was mixed with 50 mL extracting solution). The sample was filtered and potassium content was directly measured by flame AES.

K extraction by Mehlich 1 method. The diluted double acid extracting solution (Mehlich 1) contained 0.05 M HCl and 0.0125 M H₂SO₄. The soil-to-solution ratio was 1 : 5 (10 g of accurately weighted soil samples were mixed with 50 mL of extracting solution for 5 min). Potassium content in the filtered extract was determined by the flame AES.

K extraction by CaCl₂. Five grams accurately weighted soil samples were treated with 50 mL 0.01M CaCl₂ solution for 30 min [22]. The obtained extract was sent for flame AES.

K extraction by NH₄OAc/HOAc. The extracting solution 0.5 M NH₄OAc/HOAc, pH 4.5 was prepared according to [23]. Ten grams of the soil samples were mixed with 50 mL of the extracting solution for 30 min and filtered. The solution was analysed by Flame AES.

Potassium determination by Flame Atomic emission spectrometry. Potassium content in the obtained extracts was determined by Jenway PFP7 Flame photometer equipped with interference type color filter and PIN diode detector. The emitted light was measured at 766 nm. Low temperature air/propane gas mixture was used. The airflow was 6 L/min at 12 psi, sample uptake 4.5 mL/min. The optimization of flame conditions were made daily to adjust propane gas flow, applying a maximum signal for nominal K standard as a optimization criterion. The concentration of the nominal standard was chosen depending on the extracting procedure and expected sample concentrations. The same extracting solution was used to set the instrument zero. A calibration curve (readout as a function of concentration of standard solutions) was constructed for each extracting solution at the optimal conditions. During measurements, a nominal standard solution was used to monitor and fine adjusting the optimal flame conditions.

A certified reference material (**CRM**) containing 1000 mg/L (TraceCert, Lot BCBV7454- Sigma-Aldrich) was used for preparation of calibration standards by appropriate dilution with extracting solutions. A calibration

Table 1

Sources of uncertainty and expressions used for its calculation

Expressions used to calculate the uncertainty of the methods	Sources of uncertainty
$u_c(CON) = C * \sqrt{u_{rel}^2(CA) + u_{rel}^2(m) + u_{rel}^2(V)}$	$u_c(CON)$ – combined uncertainty; $u_{rel}(CA)$ – standard uncertainty associated with calculation of analyte concentration from calibration curve; $u_{rel}(m)$ – standard uncertainty associated with weighting the sample aliquot; $u_{rel}(V)$ – standard uncertainty associated with measurement of the volume of the extracting solution all presented in terms of relative uncertainty;
$u_{rel}(CA) = \sqrt{u_{rel}^2(Stand) + u_{rel}^2(Cal) + u_{rel}^2(Repet)}$	$u_{rel}(Stand)$ – standard uncertainty associated with preparation of calibration standards; $u_{rel}(Cal)$ – standard uncertainty associated with calibration curve – linear regression; $u_{rel}(Repet)$ – standard uncertainty obtained in repeatability conditions;
$u_{rel}(Stand) = \sqrt{u_{rel}^2(Prim) + u_{rel}^2(Dil)}$	$u_{rel}(Prim)$ – uncertainty derived from the preparation of the primary standard solution; $u_{rel}(Dil)$ – uncertainty derived from the preparation of the calibration curve at four concentration levels by diluting the standard solution;
$u(Cal) = \frac{1}{b} \sqrt{s_{resid}^2 * \frac{1}{n} + (Ci - Cm)^2 * s_b^2}$	u_{Cal} - uncertainty derived from linear least squares calibration;
$u(Repet) = \frac{S_s}{\sqrt{n}}$	S_s is the standard deviation from the sample replicates; n is the number of replicates of each sample when analyzed in routine analysis;
$U = k * u_c$	U is expanded uncertainty calculated at coverage factor of 2 ($k = 2$) and 95% confidence level;

curve in working range specific for each of extractants was prepared: 10-100 $\mu\text{g mL}^{-1}$ K in $\text{NH}_4\text{OAc}/\text{HOAc}$; 5-20 $\mu\text{g mL}^{-1}$ K in acetate-lactate extracting mixture, 5 - 40 $\mu\text{g mL}^{-1}$ in Mehlich 1 and 5-20 $\mu\text{g mL}^{-1}$ K in CaCl_2 . The studied extraction mixtures were used as blank samples for adjusting the instrument blank. The intensity of the emission of each solution was measured in duplicate and the mean value was used for standard curve construction. The analytical function was obtained applying linear and polynomial regression by least squares method to the obtained data.

Uncertainty and trueness estimation. Method trueness was estimated as bias and relative spike recovery by standard addition of K solution (certified reference material (CRM) 1000 $\text{mg/L} \pm 4 \text{ mg/L}$ (TraceCert, Lot BCBV7454- Sigma-Aldrich) to the soil sample before extraction. The modified acetate-lactate extraction and flame AES were applied for bias study. An aliquot of standard KCl solution was added to 2 g of dry soil, homogenized and stayed to equilibrate for 3 days. Seven replicates were made and the average concentration was used for calculation. Bias was calculated as a difference between content of potassium in the studied soil before and after standard addition. Recovery was calculated as a relative ratio between content of potassium in standard addition obtained after analysis and calculated from the volume and the concentration of standard solution of KCl. The uncertainty of bias was calculated from standard deviation of potassium in soil after standard addition (7 replicates analyzed) according to Cuadros-Rodríguez et al. [24]. Method uncertainty was estimated according to [24-26]. The sources of uncertainty and the applied equations are presented in Table 1.

Internal Quality control. The instrument performance was monitored applying the following criteria: (1) the photometer reading within 0.05% of the standard concentration of Quality control standard solution (prepared by appropriate dilution of CRM – TraceCert, Lot BCBV7454); (2) the regression coefficient for the calibration equation $R^2 \geq 0.98$; (3) the relative standard deviation of a set of three Flame Photometer readings $\leq 3 \%$. If the criteria were not met, the optimization of flame conditions and calibration were renewed.

The method and laboratory performance were monitored by analyzing test soil samples according to EURACHEM/CITAC Guide [25]. The soil samples containing: 22 and 52 $\text{mg K}_2\text{O}/100$ dry soil and used as internal quality control samples, which were analyzed with each batch of measurements. The control charts were constructed and the warning and action limits were set at $\pm 10 \%$ and $\pm 15 \%$ of the mean value, respectively.

Statistical processing of data. Each soil sample was analysed in duplicates. The mean concentration was calculated along with standard deviation at 95% confidence interval. Correlation analysis was made by Excel Data Analysis tool pack at 0.05 significance level and 22 degree of freedom.

RESULTS AND DISCUSSION

Estimation of the uncertainty of studied methods for K determination

The estimation of the uncertainty associated to analytical methods aimed at establishing the comparability of results. In this study the uncertainty was estimated following the modelling approach [24-27]. The measurand was analyte concentration in a soil sample, expressed in $\text{mg K}_2\text{O} (100\text{g})^{-1}$. The model equation was:

$$C = \frac{CA \cdot V \cdot 100 \cdot 1.205}{m \cdot 1000}, \quad (1)$$

where CA was the analyte concentration obtained from the calibration (in mg L^{-1}) and m was the sample weight (in g) and V was the extract volume (in mL), 1.205 was transfer coefficient from K to K_2O concentration. The dispersion of results around the true value depended on:

- Estimation of the analyte concentration from the calibration curve;
- Measurement of the weight of the sample;
- Measurement of volume of extracting solution.

The sources of uncertainty are presented in Table 1 and the results from uncertainty calculations in Table 2. As can be seen from the data presented in Table 2, the most influencing factors were calibration and repeatability. It should be pointed out that for more realistic estimation of repeatability the data from samples passed all the stages of analysis were taken. The standard uncertainty in repeatability conditions was estimated on the routine laboratory practice in which 2 parallel samples were analyzed ($n = 2$, Table 1). Thus, if better uncertainty was desired the efforts should be put to improve calibration procedure. The main source of errors was the deviation of calibration curve from linear fit, assessed by correlation coefficient (R^2) and standard deviation of residues (s_{reg}). Higher calibration uncertainty (Table 2) was noticed for $\text{NH}_4\text{OAc}/\text{HOAc}$ method in the concentration range 20-100 mg K/L (as presented in the Jenway Protocol [23]). The lowest calibration uncertainty was observed by ALM method: 0.17 mg L^{-1} . To improve the calibration uncertainty polynomial least squares fit was studied. The calibration equations for each of the studied extraction methods are presented in Table 3. Polynomial fit applied to the experimental data resulted in the correlation coefficient $R^2 > 0.999$ in each of the studied working ranges. The working range for a given extractant was chosen based on the expected concentration of K_2O in the studied soil samples. However, dilution of the extracts was made if the concentration of K exceeded the calibration range. In further study, the polynomial fit was applied to calculate the plant available potassium in soil extract.

Estimation of trueness

The method trueness defined as a “closeness of agreement between the average of an infinite number of replicate measured quantity values and a reference quantity value” and inversely related to systematic measurement error [28] was practically assessed by bias and relative

Table 2

Uncertainty of K determination in soils by the studied methods

Method	Sources of uncertainty	Value x	Standard Uncertainty $u(x)$	Relative Standard Uncertainty $u(x)/x$	Combined Uncertainty u_c	Expanded Uncertainty U
ALM	Linear least squares calibration	10 mg L ⁻¹	0.17 mg L ⁻¹	0.017	0.049 mg (100 g) ⁻¹	0.10 mg (100 g) ⁻¹
	Repeatability	28 mg (100 g) ⁻¹	0.84	0.03		
	Concentration of the standard solutions	10 mg L ⁻¹	0.025 mg L ⁻¹	0.0025		
	Concentration of sample in extract solution	0.04 g mL ⁻¹	0.00023 g mL ⁻¹	0.0058		
Mehlich 1	Linear least squares calibration	10 mg L ⁻¹	1.4 mg L ⁻¹	0.14	0.31 mg (100 g) ⁻¹ 0.64 mg (100 g) ⁻¹	
	Repeatability	10 mg (100 g) ⁻¹	0.73	0.073		
	Concentration of the standard stock solution	10 mg L ⁻¹	0.025 mg L ⁻¹	0.0025		
	Concentration of sample in extract solution	0.2 g mL ⁻¹	0.0023	0.012		
CaCl ₂	Linear least squares calibration	10 mg L ⁻¹	0.75 mg L ⁻¹	0.08	0.085 mg (100 g) ⁻¹ 0.17 mg (100 g) ⁻¹	
	Repeatability	5 mg (100 g) ⁻¹	0.4	0.08		
	Concentration of the standard stock solution	10 mg L ⁻¹	0.025 mg L ⁻¹	0.0025		
	Concentration of sample in extract solution	0.1 g mL ⁻¹	0.0006	0.0057		
AA	Linear least squares calibration	60 mg L ⁻¹	32 mg L ⁻¹	0.53	0.57 mg (100 g) ⁻¹ 1.1 mg (100 g) ⁻¹	
	Repeatability	28 mg (100 g) ⁻¹	0.94	0.03		
	Concentration of the standard stock solution	60 mg L ⁻¹	0.025 mg L ⁻¹	0.0025		
	Concentration of sample in extract solution	0.2 g mL ⁻¹	0.0006	0.0001		

Table 3

Calibration equations of the studied methods for determination of plant available potassium by Flame AES after different extraction procedures

Method	Concentration range (mg K/L)	N	Equation	R ²	n
ALM	5 ÷ 20	4	$y = -0.0053x^2 + 1.1255x + 1.162$	0.9999	6
Mehlich 1	5 ÷ 40	5	$y = -0.0089x^2 + 1.3795x + 0.767$	0.9996	3
AA	20 ÷ 100	5	$y = -0.0036x^2 + 1.2690x + 11.553$	0.9992	6
CaCl ₂	5 ÷ 20	4	$y = -0.0213x^2 + 1.6776x + 0.550$	0.9998	2

N – number of calibration points, n – number of calibration runs, R – correlation coefficient.

spike recovery [25, 29]. A standard addition approach was applied in two variants: (1) to assess method bias – the standard addition of potassium CRM solution before extraction and (2) to assess measurement bias – standard addition potassium CRM after extraction and before measurement [25, 29, 30]. The advantage of the applied approach was that it allowed the trueness to be evaluated in sample types usually encountered in the laboratory. It is well recognized that the efficiency of extracting solution depends not only on its composition and procedure used, but on the soil type and its physical and chemical properties [1, 31]. Thus, soil matrix varied widely from sample to sample and trueness estimation could be widely influenced. As a drawback, it should be noted that the spiked potassium could not reach full equilibrium with the soil sample and the obtained bias appeared approximate estimate of bias in available potassium in soil [25, 29, 30]. The defined volume of standard KCl solution was added to studied soils

and let to equilibrate for 3 days. The spiked soils were subjected to full extraction-determination procedure by ALM and flame AES. The obtained average bias was $1.49 \text{ mg K}_2\text{O (100 g)}^{-1}$ dry soil, corresponding to 111 % recovery. The uncertainty of associated to the recovery was $0.03 \text{ mg K}_2\text{O (100 g)}^{-1}$ calculated according to [24]. The hypothesis of no bias correction was assumed and the bias was included in the combined uncertainty: $u_c = 1.6 \text{ mg K}_2\text{O (100 g dry soil)}^{-1}$ and expanded uncertainty: $3.3 \text{ mg K}_2\text{O (100 g dry soil)}^{-1}$.

Assessment of soil characteristics

The results from analysis of 24 agricultural (arable) soil samples from two regions in Bulgaria are presented in Table 4. Active acidity (in H_2O), organic carbon content, cation exchange capacity, main nutrients content, exchangeable Na^+ , K^+ , Ca^{2+} , Mg^{2+} and Al^{3+} , as well as clay fractions were determined. The soils pH varied between 6.97 - 8.40. CEC in all studied samples was between

Table 4

Characterization of topsoil properties

Sample	pH	EC, $\mu\text{S/cm}$	CEC, cmolc/kg	Exchangeable ions, mg/kg					orgainc C, %	Ex- changea- ble PO_4^{3-} , mg/kg	ΣNO_3^- + NH_4^+ , mg/kg	Clay fraction, %
				Na^+	K^+	Ca^{2+}	Mg^{2+}	Al^{3+}				
Region 1												
1R	6.94	185	36.9	5.2	12.6	649	49.6	0.1	2.48	56	14	15
2R	7.38	172	37.7	7.4	7.7	662	53.5	0.2	2.39	74	12	18
3R	7.53	191	32.2	5.4	13.5	569	40.7	0.1	2.61	98	16	28
4R	7.52	191	39	13.9	7.6	660	70.2	0.1	3.35	44	19	15
5R	7.4	798	37.5	54.7	10.9	573	102	0.2	2.65	59	13	16
6R	7.42	1214	36.8	68	8	550	109	0.7	2.96	50	13	15
7R	7.68	296	20.1	47.3	5	317	49.5	0.1	0.98	40	5	22
8R	7.54	181	29.2	5.4	16.2	520	33.2	0.2	1.65	157	10	25
9R	6.97	105	23.5	8.5	16.2	373	53	10.8	2.69	21	26	29
10R	6.47	97	21.9	4.8	18.2	350	47	1.5	2.8	42	21	26
11R	7.21	182	33	3.9	8.6	596	35.6	<0.02	2.64	46	12	22
12R	7.67	163	35.8	3.4	7.9	649	37.6	0.3	1.93	114	6	15
max	7.7	1214	39.0	68	18.2	662	109	10.8	3.35	157	26	29
min	6.5	97	20.1	3.4	5	317	33.2	0.1	0.98	21	5	15
median	7.4	183	34.4	6.4	9.75	571	49.55	0.2	2.63	53	13	20
mean	7.3	314	32.0	19.0	11.0	539	56.7	1.3	2.43	67	14	21
N	12	12	12	12	12	12	12	12	12	12	12	12
Region 2												
1F	7.49	151	24.6	2.3	13.8	466	10.5	<0.05	2.8	92	13	29
2F	7.84	133	24.4	3.1	15.4	463	10.7	<0.05	2.17	95	12	21
3F	8.05	169	20.5	3.4	24.7	380	10.8	<0.05	2.56	137	13	21
4F	7.88	166	25.4	3.9	24.1	476	12.4	<0.05	3.02	95	7	39
5F	7.42	147	24.7	2.5	17.1	470	9.4	<0.05	2.87	97	16	38
6F	7.57	308	24.9	2.7	12.5	474	9.8	<0.05	2.66	70	21	48
7F	7.67	140	17.4	3	12.1	333	4.6	<0.05	4.04	163	12	56

8F	7.76	129	17.8	2.3	13.8	340	5.3	<0.05	3.9	87	16	39
9F	7.75	143	21.8	2.5	19.5	417	6	1.3	4.12	133	17	27
10F	7.89	195	32.1	3	10.8	623	8.5	0.5	4.02	36	17	20
11F	8.26	162	28.5	2.4	12.4	543	12.7	0.1	1.61	175	10	41
12F	8.4	177	34.5	2.6	20	654	15.5	0.5	1.65	586	13	35
max	8.4	308	34.5	3.9	24.7	654	15.5	1.3	4.12	586	21	56
min	7.4	129	17.4	2.3	10.8	333	4.6	0.1	1.61	36	7	20
median	7.8	156	24.7	2.65	14.6	468	10.2	0.5	2.84	96	13	37
mean	7.8	168	24.7	2.8	16.4	469.9	9.7	0.6	2.95	147	14	35
N	12	12	12	12	12	12	12	12	12	12	12	12
Region 1 and 2												
max	8.4	1214	39	68	24.7	662	109	10.8	4.12	586	26	56
min	6.5	97.2	17.4	2.3	5	317	4.6	0.1	0.98	21	5	15
median	7.6	170.4	27.0	3.7	13.1	498	24.4	0.2	2.66	90	13	26
mean	7.6	241.3	28.3	10.9	13.7	504	33.2	1.1	2.69	107	14	28
N	24	24	24	24	24	24	24	24	24.00	24	24	24

18 - 39 cmol kg⁻¹. Organic carbon varied between 1.6 and 4.2%. The studied soils presented different fertility indices according to the classification based on the modified acetate-lactate method and accepted in Bulgaria: 30 % very high, 33 % high, 33 % medium and 4 % very low. The studied soils were chosen, from one hand, to present different soil characteristics, and from other – to study the performance of double acid method (Mehlich 1) in neutral or alkaline samples with CEC > 10 cmol_c kg⁻¹. The interest was provoked by the contradiction in the published data. First, Mehlich 1 method was designed for acidic soils with low CEC from Florida, USA (pH < 6.5 and CEC < 10 mgeq kg⁻¹) [32]. However, Hosseinpour and Samavati [12] found that Mehlich 1 well correlated with other methods for calcareous soils from Hamedan region, Iran. Considering the fact, that Mehlich 1 was fast, cheap and easy to apply we were interested to study its effectiveness on selected Bulgarian arable soils.

Comparison of extractable potassium by the studied methods

The plant available K estimated by Mehlich 1, a modified acetate-lactate procedure (ALM), NH₄OAc/HOAc (AA), CaCl₂, BaCl₂ and Olsen methods were studied. The results for extractable potassium in soil, expressed as mg K₂O/100 g dry soil, and the fertility status of the studied soils are presented in Table 5 and Table 6. A comparison between the amount of extractable potassium is illustrated on Figure. The results from Olsen method (0.5 M NaHCO₃) were not presented as the obtained extracts were not suitable for Flame AES determination of potassium due to some interferences.

CEC was determined by Hendershot and Duquette method [21] after extraction with BaCl₂ and ICP-AES determination, thus providing data about potassium in extract. As can be seen from the data, BaCl₂ extracted more potassium than CaCl₂ due to the higher concentration

of Ba²⁺ and 1 : 10 soil-to-liquid ratio, as well as to higher sensitivity of ICP-OES measurement.

Two soil-to-liquid ratios were studied with Mehlich 1 reagent. It was noticed that in calcareous soils with pH > 7.5 the acidity of extractant was lowered due to neutralization with carbonates. The pH of extract was higher than 4.5 and reached 7 in some samples. In contrast, extract of soils samples with pH < 7 kept the original level of acidity of Mehlich 1 extractant (pH < 2.5). Thus, Mehlich 1 reagent did not guarantee the same conditions of extraction in different types of soils, which could affect the parts of extracted potassium. In attempt to overcome changes in extraction conditions, two series of experiments were conducted: (1) increasing soil-to-liquid ratio up to 1 : 10 with the same concentration of acids and (2) doubled concentration of Mehlich 1 extractant with the same soil-to-liquid ratio (1 : 5). Increased soil-to-liquid ratio (1 : 10) preserved the pH of extractant in the majority of calcareous soils, the obtained extract was clear and suitable not only for K determination by Flame AES, but for spectrophotometric determination of other nutrients. In the second series of experiment, the increased concentration of acids in the extraction mixture resulted in unchanged pH, but the obtained solution was unclear and needed additional treatment with charcoal before the measurement. As the results showed, the increased soil-to-liquid ratio was better approach to keep extraction conditions and to obtain the same part of available potassium extracted. It was in contrast to observations made by Quin *et al.* [33] that the increased soil-extractant ratio did not improve the results.

The results in Table 6 and Figure demonstrated that the amount of total and extractable K varied widely with the method used. It was shown that in the same soil sample, the content of exchanging potassium species varied greatly when using different extraction methods: from 4.3 to 10.1 times (Region 1) and from 3.2 to 10.9 times (Region 2). Of the tested methods, ALM extracted

Table 5

Extractable potassium determined by different extractants and Flame AES and corresponding fertility indices (^a [1]; ^b [35]); ^c [23]; ^d [4]). Standard deviations of 2 parallel samples are presented in the brackets next to the mean value

Sample	K ₂ O, mg (100g) ⁻¹								
	ALM	Fertility ^a	Mehlich 1	Fertility ^b	CaCl ₂	Fertility ^c	NH ₄ Ac/HAc	Fertility ^d	BaCl ₂
Region 1									
1R	38.8 _(0.2)	very high	8.0 _(1.3)	high	5.1 _(0.1)	low	31.1 _(0.86)	very low	15.2 _(0.1)
2R	22.8 _(0.5)	medium	3.4 _(0.3)	low	2.5 _(0.1)	low	22.3 _(0.8)	very low	9.3 _(0.4)
3R	50.3 _(0.9)	very high	7.7 _(1.8)	high	6.5 _(0.04)	low	39.8 _(0.2)	low	16.3 _(0.3)
4R	24.7 _(0.2)	medium	6.3 _(0.7)	medium	3.0 _(0.1)	low	23.3 _(1.0)	very low	9.2 _(0.8)
5R	36.3 _(0.5)	high	6.6 _(0.1)	medium	4.1 _(0.2)	low	32.0 _(0.7)	very low	13.2 _(1.1)
6R	28.2 _(0.4)	medium	6.1 _(1.0)	medium	2.8 _(0.1)	low	23.7 _(0.8)	very low	9.7 _(0.1)
7R	14.1 _(0.6)	low	1.8 _(0.7)	low	1.5 _(0.6)	low	13.1 _(0.7)	very low	6.1 _(0.7)
8R	43.1 _(1.0)	very high	7.2 _(3.3)	high	5.4 _(0.04)	low	35.7 _(0.8)	very low	19.5 _(1.4)
9R	36.0 _(1.5)	high	13.7 _(1.9)	very high	5.7 _(0.3)	low	27.5 _(0.6)	very low	16.0 _(0.2)
10R	43.7 _(1.2)	very high	17.1 _(2.1)	very high	10.1 _(1.3)	low	43.5 _(2.4)	low	15.6 _(0.3)
11R	27.0 _(1.6)	medium	3.9 _(2.0)	medium	2.7 _(0.2)	low	24.5 _(0.8)	very low	10.4 _(0.6)
12R	30.2 _(0.9)	high	3.5 _(0.7)	medium	3.3 _(0.2)	low	26.6 _(0.1)	very low	9.5 _(0.4)
max	50.3		17.1		10.1		43.5		19.5
min	14.1		1.8		1.5		13.1		6.1
median	33.1		6.5		3.7		27.0		11.8
mean	32.9		7.1		4.4		28.6		12.5
N	12		12		12		12		12
Region 2									
1F	38.8 _(0.4)	high	7.2 _(0.7)	high	5.3 _(0.1)	low	26.3 _(1.0)	very low	14.7 _(0.1)
2F	31.5 _(0.5)	high	6.0 _(0.9)	medium	6.4 _(0.03)	low	26.5 _(1.0)	very low	8.6 _(0.8)
3F	57.7 _(0.3)	very high	13.7 _(1.4)	very high	14.0 _(1.2)	low	45.3 _(1.0)	low	15.6 _(0.3)
4F	49.5 _(0.1)	very high	9.9 _(2.2)	high	10.0 _(1.3)	low	29.4 _(1.0)	very low	9.1 _(0.9)
5F	39.5 _(0.5)	high	6.5 _(2.3)	medium	7.4 _(1.0)	low	29.8 _(1.0)	very low	12.9 _(0.2)
6F	27.5 _(0.7)	medium	5.5 _(0.3)	medium	5.0 _(0.1)	low	23.8 _(0.8)	very low	8.7 _(0.4)
7F	43.2 _(0.3)	very high	7.9 _(1.9)	high	5.4 _(0.1)	low	20.3 _(0.4)	very low	9.7 _(0.1)
8F	32.3 _(0.5)	medium	6.0 _(0.9)	medium	6.3 _(0.7)	low	23.1 _(0.8)	very low	9.0 _(0.9)
9F	26.1 _(0.1)	medium	2.7 _(1.7)	low	8.4 _(0.1)	low	27.4 _(0.6)	very low	22.3 _(1.0)
10F	28.0 _(0.9)	medium	7.5 _(1.7)	high	2.1 _(0.1)	low	22.9 _(1.0)	very low	10.4 _(0.6)
11F	17.4 _(0.3)	medium	3.5 _(1.2)	medium	3.4 _(0.2)	low	26.7 _(0.1)	very low	6.3 _(0.7)
12F	25.9 _(0.7)	high	6.3 _(0.4)	medium	6.0 _(0.8)	low	26.5 _(1.0)	very low	19.8 _(1.3)
max	57.7		13.7		14.0		45.3		22.3
min	17.4		2.7		2.1		20.3		6.3
median	31.9		6.4		6.2		26.5		10.1
mean	34.8		6.9		6.6		27.3		12.2
N	12		12		12		12		12
Region 1 and 2									
max	57.7		17.1		14.0		45.3		22.3
min	14.1		1.8		1.5		13.1		6.1
median	31.9		6.4		5.4		26.5		10.4
mean	33.9		7.0		5.5		28.0		12.4
N	24		24		24		24		24

the highest amount of K (mean 33.4 mg K₂O/ 100 g, median 31.9 mg K₂O/ 100g, ranged 14.1- 57.7 mg K₂O/ 100g) and CaCl₂ removed the lowest amount of K (mean 5.6 mg K₂O/ 100g, median 5.4 mg K₂O/ 100g,

ranged 1.5-14.0 mg K₂O/ 100g). On average ALM extracted 1.2, 3.1, 4.8, and 5.9 times more K than did AA, BaCl₂, Mehlich 1 and CaCl₂, respectively. As can be seen ALM and AA extracted almost the same amount of K

Table 6

Soil potassium fertility indices applied in this study

Method	K ₂ O mg/100g	Fertility index	References
ALM	<12	very low	[1]
	13-16	low	
	17-28	medium	
	29-38	high	
	>38	very high	
Mehlich 1	<1.5	very low	[34]
	1.5-3.5	low	
	3.5-7	medium	
	7-13	high	
	>13	very high	
NH ₄ OAc(HOAc)	0-36	Very low	[23]
	37-72	Low	
	73-145	Medium	
	146-241	High	
	>242	Very high	
CaCl ₂	<15	Low	[4]
	15-25	Medium	
	25-80	High	
	>80	Excessive	

(Table 6). The possible explanation of this observation could be found in similarity of composition and acidity of two extracting mixtures and probably the mechanism of extraction [12]. Both methods extracted the same part of exchangeable and non-exchangeable soil potassium. The higher quantities of ALM extracted potassium could be due not only to the extractant composition, but to the highest soil-to-liquid ratio and longest extraction time. The lower quantity of extracted potassium by Mehlich 1 solution could be explained by the high pH and CEC of the studied soils, which lower the extraction capacity of Mehlich 1 solution by neutralization.

Table 7

Correlation analysis (P<0.05; df=22; 95% confidence level)

	ALM	Mehlich1	CaCl ₂	NH ₄ OAc/HOAc
ALM				
Mehlich1	0.7956 ^a			
CaCl ₂	0.7688 ^a	0.7968 ^a		
NH ₄ OAc/HOAc	0.9494 ^a	0.6816 ^a	0.7352 ^a	
BaCl ₂	0.5662 ^a	0.5803 ^a	0.4385 ^b	0.6629 ^a

a – significant P < 0.05, b – insignificant P > 0.05.

It was well recognized that the amount of nutrients extracted by a given extraction procedure was not a direct measure of plant availability of nutrients but an index correlated to field calibration [1, 34]. The relationships between K-extracted by different soil testing extractants could provide an information about the parts of extractable potassium and efficiency of given extractant for specific soil types. The correlation coefficients between K extracted by four studied extracting solutions are shown in Table 7. There was a positive correlation (at 0.05 level of significance) between the quantity of extractable potassium by the studied extraction methods, except between the results from CaCl₂ and BaCl₂. It should be pointed out that CaCl₂ and BaCl₂ based extractions differed not only in extracting procedure, but also in detection technique. Although, the results demonstrated that the concentrations of extractable K varied widely with the method used (Table 5), a significant correlation between the obtained results was observed (Table 7). It could be hypothesized that each extracting procedure desorbed different amount of K, but the similar K fractions were determined [12]. However, due to the variety of applied extractants and mechanism of extractions comparison of fertility indices based on different test methods could provoke unrealistic estimation of potassium soil status and field calibration needed for Bulgarian arable soils.

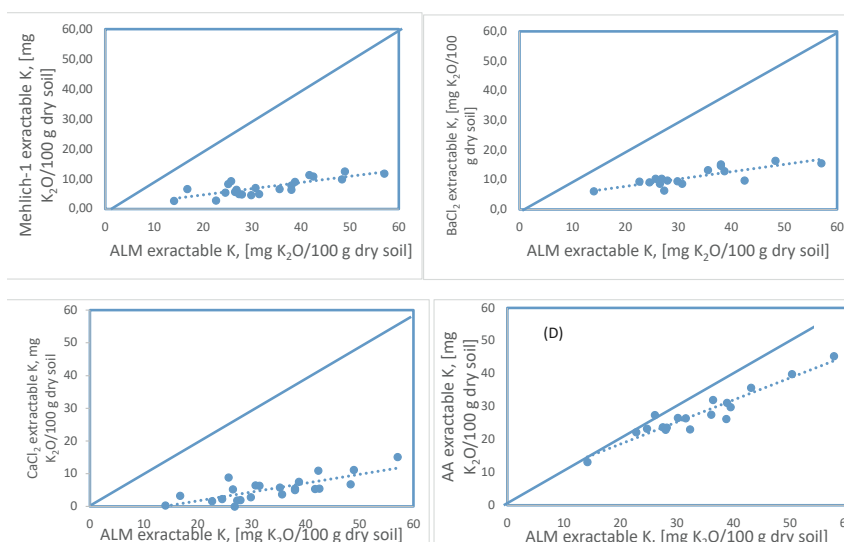


Figure. Comparison of extractability of various soil test methods for K from soil samples collected at a sample depth of 0 to 30 cm from agricultural fields in Bulgaria. The dotted line represents the linear regression fit and a solid line - a 1:1 relationship between the tests being compared.

CONCLUSIONS

A comparative study on four plant available potassium methods was conducted for Bulgarian arable soils. A modified acetate-lactate method (ALM), Mehlich 1, $\text{NH}_4\text{OAc}/\text{HOAc}$ (AA), CaCl_2 and BaCl_2 were studied. The results showed that extractable potassium lowered in the following order $K_{\text{ALM}} \geq K_{\text{AA}} > K_{\text{Mehlich1}} > K_{\text{BaCl}_2} > K_{\text{CaCl}_2}$. The amount of extractable K varied widely with the method used with ALM method extracting the highest amount of exchangeable K. A significant correlation between the obtained concentrations of extractable K by the studied methods was observed, indicating that similar fractions of K was extracted. The obtained results provided a base for further study on correlation between extractable K and soil fertility indices for particular soil types and climatic regions in Bulgaria.

ACKNOWLEDGEMENT

The authors of this study were thankful to prof. Margarita Nikolova for the idea and discussions; to Mr. Vencislav Nekov (Sembodja, Ltd., Sofia, Bulgaria) for providing soil samples. Lyudmila Angelova acknowledged the financial support of the National Program for Support of Young Researchers 2018–2020 funded by the Bulgarian Ministry of Education and Science. The authors are thankful to Ms. Yulia Ivanova for making some of the laboratory experiments.

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